

Topic 1 Reaction Rates

1. Which one of the following will not effect the rate of reaction if altered?
 - A concentration of solution
 - B temperature of reactants
 - C particle size of reactants
 - D volume of solution

2. In order for a chemical reaction to take place particles must
 - A not collide
 - B collide with enough energy only
 - C collide with the correct orientation only
 - D collide with enough energy and the correct orientation

3. When the concentration of a reactant increases the reaction rate
 - A increases
 - B decreases
 - C stays the same

4. When the particle size of a reactant is reduced the reaction rate

- A increases
- B decreases
- C stays the same

5. When the temperature is increased the reaction rate

- A increases
- B decreases
- C stays the same

6. When the concentration of a reactant increases the reactant particles

- A only collide more frequently
- B only collide with more energy
- C collide with both more energy and more frequently

7. When the temperature of reactants is increased the reactant particles

- A only collide more frequently
- B only collide with more energy
- C collide with both more energy and more frequently

8. When the particle size of a reactant is decreased the reactant particles

- A only collide more frequently
- B only collide with more energy
- C collide with both more energy and more frequently

9. If you reduce the particle size of a reactant then

- A the surface areas is reduced
- B the surface area is increased
- C there is no change in surface area

10. The average rate of reaction can be calculated by

- A measuring the change in mass only
- B measuring the volume of a product gas and measuring time
- C measuring the volume of a product gas only