

Ionic Formula Practice Homework

1. What is the ionic formula for magnesium oxide

- A MgO
- B $\text{Mg}^{2+}\text{O}^{2-}$
- C $\text{Mg}^{3+}\text{O}^{2-}$
- D $\text{Mg}^{2+}\text{O}^{2-}_2$

2. What is the ionic formula for sodium oxide

- A $\text{Na}^{3+}\text{O}^{2-}$
- B $\text{Na}^{2+}\text{O}^{2-}$
- C $\text{Na}^+_2\text{O}^{2-}$
- D $\text{Na}^+\text{O}^{2-}_2$

3. What is the ionic formula for aluminium oxide

- A $\text{Al}^{3+}\text{O}^{2-}_2$
- B $\text{Al}^{3+}_3\text{O}^{2-}_2$
- C $\text{Al}^{2+}_2\text{O}^{3-}_3$
- D $\text{Al}^{3+}_2\text{O}^{2-}_3$

4. What is the ionic formula for calcium hydroxide

- A $\text{Ca}^{3+}_2\text{OH}^{2-}_3$
- B $\text{Ca}^{2+}(\text{OH}^-)_2$
- C $\text{Ca}^{3+}_2\text{OH}^{2-}_3$
- D $\text{Ca}^{3+}(\text{OH}^-)_3$

5. What is the ionic formula for lithium hydroxide

- A $\text{Li}^+\text{OH}^{2-}$
- B $\text{Li}^+_2\text{OH}^{2-}$
- C Li^+OH^-
- D $\text{Li}^+(\text{OH}^-)_2$

6. What is the ionic formula for aluminium hydroxide

- A $\text{Al}^{3+}(\text{OH}^-)_3$
- B $\text{Al}^{3+}_3(\text{OH}^-)$
- C $\text{Al}^{3+}_2\text{OH}^-$
- D $\text{Al}^{3+}_2\text{OH}^{2-}_3$

7. What is the ionic formula for potassium carbonate

- A $\text{K}^{3+}_2\text{CO}^{2-}_3$
- B $\text{K}^+_2\text{CO}^{2-}_3$
- C $\text{K}^{2+}_2\text{CO}^{2-}_3$
- D $\text{K}^+\text{CO}^{2-}_2$

8. What is the ionic formula for calcium carbonate

- A $\text{Ca}^{2+}_2\text{CO}^{2-}_3$
- B $\text{Ca}^{2+}_3\text{CO}^{2-}_2$
- C $\text{Ca}^{2+}\text{CO}^{2-}_3$
- D $\text{Ca}^{2+}_2(\text{CO}^{2-}_3)_3$

9. What is the ionic formula for aluminium carbonate

- A $\text{Al}^{3+}\text{CO}_3^{2-}$
- B $\text{Al}^{3+}(\text{CO}_3^{2-})_2$
- C $\text{Al}_2^{3+}(\text{CO}_3^{2-})_3$
- D $\text{Al}_3^{3+}(\text{CO}_3^{2-})_2$

10. What is the ionic formula for ammonium carbonate

- A $\text{NH}_4^+\text{CO}_3^{2+}$
- B $(\text{NH}_4^+)_2\text{CO}_3^{2-}$
- C $(\text{NH}_4^+)_3\text{CO}_3^{2-}$
- D $\text{NH}_4^{2+}\text{CO}_3^{3-}$